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(54) Title: ISOMERIZATION CATALYST AND PROCESS FOR ITS USE			
(57) Abstract <p>The present invention is a process and a catalyst for isomerizing normal and slightly branched C₄ to C₇ hydrocarbons. The catalyst comprises a Group VIII metal on Beta zeolite.</p>			

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ISOMERIZATION CATALYST AND PROCESS FOR ITS USE

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FIELD OF THE INVENTION

The present invention is a process for alkane isomerization. More specifically, the present invention is a process for isomerizing C₄ to C₇ paraffins to increase their octane rating.

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BACKGROUND OF THE INVENTION

The petroleum distillate fraction that contains C₄ to C₇ hydrocarbons is relatively low in octane because it contains substantial amounts of low octane, normal paraffins. For example, normal C₅ has a blending RON of 62 and normal C₆ has a blending RON of 19 (blending RON will be hereinafter called "RON"). However, when these paraffins are isomerized to form branched paraffins, their RON increases dramatically. For example, isopentane (2-methylbutane) has a RON of 99 and isohexane (2-methylpentane) has a RON of 83. Generally, the RON will increase with even higher branching (e.g., 2,2-dimethylbutane has a RON of 89).

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Several catalysts have been used to isomerize these lower octane paraffins into the branched, higher octane paraffins. Examples are shown in U.S. Patent No. 4,374,296 issued February 15, 1983 to Haag et al.; U.S. Patent No. 3,432,568 issued March 11, 1969 to Miale et al.; U.S. Patent No. 3,673,267 issued June 27, 1972 to Chen et al.; and U.S. Patent No. 4,665,272 issued May 12, 1987 to Bakas et al. Haag et al. disclose isomerizing paraffins using intermediate pore zeolites, which have a constraint index between 1 and 12. Miale et al. describe hydroisomerizing saturated aliphatic and cyclic hydrocarbons by contacting them with a dual functional catalyst comprising mordenite and a catalytic metal. Chen et al. describe a process for isomerizing paraffins using mordenite having a silica to alumina ratio between 20:1 and 60:1. Bakas et al. disclose isomerizing paraffins using a crystalline aluminosilicate having a catalytic metal.

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However, even though these prior art catalysts are useful, there remains a need for a new catalyst that is: (1) highly active; (2) highly selective for producing high octane liquid product; and (3) sulfur tolerant. That need is satisfied by the invention that is detailed below.

SUMMARY OF THE INVENTION

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According to the present invention, a catalyst and a process for using the catalyst are provided for isomerizing C₄ to C₇ hydrocarbons. The catalyst comprises a Group VIII metal and Beta zeolite and the process comprises contacting the catalyst with a feed having normal and branched C₄ to C₇ hydrocarbons under isomerization conditions. Preferably, the Group VIII metal is platinum and the feed substantially comprises normal and singly branched lower octane C₅ and C₆ hydrocarbons. To improve catalytic activity the catalyst is preferably calcined in a steam/air mixture at an elevated temperature after impregnation with the Group VIII metal.

Among other factors, the present invention is based on my finding that a catalyst comprising Beta zeolite and a Group VIII metal is highly active and highly selective for isomerizing C₅ and C₆ hydrocarbons. The catalyst is also surprisingly selective for isomerizing C₇ hydrocarbons. Additionally, the catalyst is relatively sulfur tolerant. Furthermore, when the catalyst is calcined in a steam/air mixture, higher activity and selectivity are achieved.

More specifically, the process for alkane isomerization comprises contacting a feed containing low octane normal and singly branched C₅ and C₆ hydrocarbons and less than 0.1 ppm sulfur with a catalyst which comprises Beta zeolite and between 0.1% and 1.0% platinum, at a temperature between 400°F and 600°F, at a pressure between 100 and 500 psig, a H₂/HC ratio between 1 and 8, and between 1 and 4 LHSV.

DETAILED DESCRIPTION OF THE INVENTION

The following discussion presents a more

detailed discussion of the present catalyst.

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The Catalyst

05 Beta zeolite is a large pore zeolite. Large pore zeolites typically have effective pore diameters between 6 and 8 Angstroms (Å). Beta zeolite has a pore opening of approximately 7 Å. Although its present structure is unknown, it is believed to have pore windows formed by twelve-membered rings of silicon and aluminum atoms.

10 Beta zeolite is substantially shown and described in U.S. Patent No. 3,308,069 and Re 28,341 to Wadlinger et al., and Breck, "Zeolite Molecular Sieves", p. 309 (1984), which are all hereby incorporated by reference in their entireties. These references describe a process for making Beta zeolite, its elemental composition and formula.

15 Preferably, the Beta zeolite of the present invention has a silica to alumina molar ratio between 10 and 1000, more preferably between 15 and 100. It is also preferable that the catalyst have high acidity. The catalyst's acidity can be increased by ammonium ion exchange and subsequent calcination to yield the hydrogen form. In addition to the hydrogen form, other forms of the zeolite can be used if they are sufficiently acidic. Preferably, 20 the catalyst is substantially free of alkali or alkaline earth metals.

25 Preferably, catalysts of the present invention contain one or more Group VIII metals, e.g., nickel, ruthenium, rhodium, palladium, iridium or platinum. The preferred Group VIII metals are iridium, palladium, and most preferably platinum. The preferred percentage of the Group VIII metal in the catalyst is between 0.05% and 5%, more preferably between 0.1% and 2.5%, most preferably between 0.1% and 1.0%.

30 Group VIII metals are preferably introduced into Beta zeolite by impregnation, occlusion, or exchange in an

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aqueous solution of an appropriate salt. When it is desired to introduce two Group VIII metals into Beta zeolite, the operation may be carried out simultaneously or sequentially. Preferably, the Group VIII metal is highly dispersed within, and on, the zeolite.

By way of example, platinum can be introduced by impregnation with an aqueous solution of tetraammineplatinum (II) nitrate, tetraammineplatinum (II) hydroxide, dinitrodiamino-platinum or tetraammineplatinum (II) chloride. In an ion exchange process, platinum can be introduced by using cationic platinum complexes such as tetraammineplatinum (II) nitrate. On the other hand, when platinum is introduced into the zeolite by occlusion a platinum complex is preferably introduced into the zeolite during its synthesis.

After the desired metal or metals have been introduced, the catalyst is preferably calcined in air, or air diluted with an inert gas. More preferably, the catalyst is calcined in a steam/air mixture at an elevated temperature after platinum impregnation.

Promoter metals can be added to the catalyst in the manner described above. These metals are preferably selected from Groups VIII, IVA, IB or VIB.

The zeolite is preferably bound with a porous matrix. The term "porous matrix" includes inorganic compositions with which the zeolite will effectively bond after appropriate thermal treatment. The matrix porosity can either be inherent or it can be introduced by chemical means. Satisfactory matrices include pumice, firebrick, diatomaceous earth and inorganic oxides. Preferred inorganic oxides include alumina, silica, naturally occurring or synthetic clays (for example, bentonite, kaolin, sepiolite, attapulgite, and halloysite). Silica or alumina are especially preferred.

Compositing the zeolite with an inorganic oxide matrix can be achieved by any suitable known method wherein the zeolite is intimately admixed with the oxide while the latter is in a hydrous state (for example, as a

5 hydrous salt, hydrogel, wet gelatinous precipitate, or in a dried state, or combinations thereof). A convenient method is to prepare a hydrous mono or plural oxide gel or cogel using an aqueous solution of a salt or mixture of salts (for example, sodium silicate). Ammonium hydroxide (or a similar base) is added to the solution in an amount sufficient to precipitate the oxides in hydrous form.

10 Then, the precipitate is washed to remove most of the water soluble salts and it is thoroughly admixed with the zeolite which is in a finely divided state. Water or lubricating agents can be added in an amount sufficient to facilitate shaping of the mix (e.g., by extrusion).

15 The Process

20 Once the catalyst has been formed it can be employed in any of the conventional types of process equipment known to the art. It may be employed in the form of pills, pellets, granules, broken fragments, or various special shapes. It can be disposed as a fixed or moving bed within a reaction zone, and the charging stock can be passed therethrough as a liquid, vapor, or mixed phase, and in upward, downward, or radial flow.

25 Alternatively, it could be prepared for use in fluidized-solid processes, in which the charging stock is passed upward through a turbulent bed of finely divided catalyst. However, in view of the danger of attrition losses of the valuable catalyst and the well-known advantages of plug flow reactors either a fixed bed system or a dense-phase moving bed system is preferred. In a fixed bed system, the feed is preheated (by any suitable heating means) to the desired reaction temperature and then passed into a isomerization zone containing the catalyst. This isomerization zone may be one or more separate reactors with suitable means to maintain the desired temperature at the entrance to each reactor. After reaction, the products from any of the foregoing process variants are separated from the catalyst, reduced to atmospheric pressure, and fractionated to recover the various components thereof.

30 Additionally, the process can be operated in once-through

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or, preferably, with recycle of normal paraffins. Recycle operations are within the knowledge of skilled practitioners.

05 As discussed above, the petroleum distillate fraction preferably contains normal and singly branched C₄ to C₇ Paraffins. The feed preferably contains few multiply branched components such as 2,2-dimethylbutane 10 and 2,3-dimethylbutane. However, when the low octane paraffins are isomerized to form multiply branched paraffins (such as the dimethylbutanes, for example), the octane number increases dramatically. This is the function of the present catalyst.

15 The present catalyst is highly active and highly selective for isomerizing C₄ to C₇ hydrocarbons. This high selectivity means that in comparison to prior art catalysts, such as mordenite, a higher product yield can be achieved when the two catalysts are run to a given octane. 20 It also means that a higher octane can be achieved when the two catalysts are run to produce a given liquid yield.

25 The present process comprises contacting the isomerization catalyst with a hydrocarbon feed under isomerization conditions. The feed is preferably a light straight run fraction, boiling within the range of 30°F to 250°F and preferably from 60°F to 200°F. Preferably, the hydrocarbon feed for the process comprises a substantial amount of C₄ to C₇ normal and singly branched hydrocarbons, more preferably low octane C₅ and C₆ hydrocarbons.

30 The pressure in the process is preferably between 50 psig and 1000 psig, more preferably between 100 and 500 psig. The liquid hourly space velocity (LHSV) is preferably between about 1 to about 10 with a value in the range of about 1 to about 4 being more preferred. It is 35 also preferable to carry out the isomerization reaction in the presence of hydrogen. Preferably, hydrogen is added to give a hydrogen to hydrocarbon ratio (H₂/HC) of between 0.5 and 10, more preferably between 1 and 8. The temperature is preferably between about 200°F and about 1000°F, 40 more preferably between 400°F and 600°F. As is well known

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to those skilled in the isomerization art, the initial selection of the temperature within this broad range is made primarily as a function of the desired conversion level considering the characteristics of the feed and of the catalyst. Thereafter, to provide a relatively constant value for conversion, the temperature may have to be slowly increased during the run to compensate for any deactivation that occurs.

10 The present catalyst is relatively sulfur tolerant. Nonetheless, a low sulfur feed is especially preferred in the present process. The feed preferably contains less than 10 ppm, more preferably less than 1 ppm, and most preferably less than 0.1 ppm sulfur. In 15 the case of a feed which is not already low in sulfur, acceptable levels can be reached by hydrotreating the feed in a pretreatment zone with a hydrotreating catalyst which is resistant to sulfur poisoning. Sulfur can be subsequently removed as hydrogen sulfide in the gas phase after 20 condensation of the liquid product. An example of a suitable catalyst for this hydrodesulfurization process is an alumina-containing support and a minor catalytic proportion of molybdenum oxide, cobalt oxide and/or nickel oxide. Hydrodesulfurization is typically conducted at 25 315°C to 455°C, at 200 to 2000 psig, and at a liquid hourly space velocity of 1 to 5.

30 It is also preferable to limit the nitrogen level and the water content of the feed. Catalysts and processes which are suitable for these purposes are known to those skilled in the art.

35 After a period of operation the catalyst can become deactivated by sulfur or coke. Sulfur and coke can be removed from the catalyst using procedures that are known to those skilled in the art.

40 The present invention will be more fully understood by reference to the following examples. They are intended to be purely exemplary and are not intended to limit the scope of the invention in any way.

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EXAMPLESExample 1

05 In a 2-liter glass liner, 8.29 gms of Mallinkrodt sodium aluminate ($\text{Na}_2\text{O} \cdot \text{Al}_2\text{O}_3 \cdot 3\text{H}_2\text{O}$) was dissolved in 259.6 gms of an aqueous solution containing 20% tetraethyl ammonium hydroxide (Aldrich). 207.5 Grams of Ludox AS-30 colloidal silica was added in a thin stream while the 10 solution was vigorously stirred. The liner was placed in a 2-liter Parr stainless steel bomb and statically heated for 6 days at 150°C. The solid, precipitate product was washed and dried, after which it was determined to be Beta zeolite by X-ray diffraction.

15 Beta zeolite was made more catalytically active by converting it into its hydrogen (proton) form as follows. Beta zeolite was layered in a thin bed and incrementally heated to 1000°F at a rate of 150°/hr. The temperature was maintained at 1000°F for 5 hours and then 20 it was increased to 1100°F for 4 more. The calcined zeolite was ion-exchanged by 4 successive treatments with NH_4NO_3 . Each treatment used the same mass of NH_4NO_3 to zeolite (a slurry of about 50 gr. zeolite/liter H_2O) and the solution was refluxed for at least 2 hours. The 25 sodium content of the zeolite was low after the fourth exchange. The zeolite was filtered, washed, dried and recalcined as above while omitting the 1100°F heating step.

30 The final Beta zeolite was impregnated with enough tetraamine platinum nitrate to achieve 0.8 wt.% platinum. Thereafter, the impregnated catalyst was dried at 250°F for approximately 14 hours then calcined at 500°F for 3 hours, and then reduced.

Example 2

35 The catalyst of Example 1 was used to isomerize a feed which substantially comprised normal and branched C_5 and C_6 hydrocarbons. The isomerization conditions were 1 LHSV, 200 psig, 6 H_2/HC , and 520°F. The feed composition and the isomerization results are shown in Table I.

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TABLE I

	<u>Compound (Wt.%)</u>	<u>Feed</u>	<u>Isomerate</u>
05	Methane	0.00	0.00
	Ethane	0.00	0.00
	Propane	0.00	0.71
	Isobutane	0.04	1.84
	N-butane	0.28	0.64
10	Iso-C ₅	12.03	21.45
	N-pentane	18.93	11.37
	2,2-DMB	0.58	9.56
	Cyclo-C ₅	4.26	3.69
	2,3-DMB	2.26	4.44
	2MP	12.55	15.36
	3MP	8.19	10.11
15	N-hexane	19.74	8.46
	MCP	15.04	9.82
	Benzene	3.75	0.00
	CHX	1.89	2.56
	C ₇ ⁺	0.45	0.00
	C ₅ + Wt.%	99.68	96.81
20	LV%	100.00	99.01
	RON	74.50	80.08

Example 3

A commercial mordenite (Zeolon 100H) was obtained from Norton Chemical Company and platinum impregnated as in Example 1. The final catalyst contained 0.8 wt.% platinum.

Example 4

This example compares the liquid yields achievable with the catalysts of Example 1 and Example 3, which both contained 0.8% platinum.

Both catalysts were used to isomerize normal C₅, C₆, and C₇ pure component feeds at 500°F, 100 psig, and 6H₂/HC. The liquid yields and corresponding octanes that were obtained for the two catalysts are shown in FIG. 1.

Example 5

Platinum was impregnated onto samples of Beta zeolite and mordenite at varying levels. Both catalysts were used to isomerize a normal hexane feed at 510°F, 100 psig, 3 LHSV, and 6H₂/HC. The results are shown in

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05 Table II. The relative intrinsic activity for each catalyst at low conversion was calculated and is shown in FIG. 2.

TABLE II

10	Catalyst Support	Pt	nC ₆ Conv.	LV%	Isomerate RON
10	H-Beta	2.4	84	94	76.0
		0.8	83	97	76.1
		0.05	82	98	75.3
15	H-Mordenite	2.4	80	92	74.1
		0.8	82	95	75.1
		0.05*	76	90	64.5

*Temp. raised to 540°F due to low activity.

Example 6

20 This example demonstrates the effect of sulfur on platinum/Beta zeolite and platinum/mordenite catalysts. Each catalyst was contacted with the feed of Table I at 520°F, 200 psig, 6H₂HC, and 1 LHSV for 24 hours. Then H₂S was continuously added to the reactor inlet H₂ and into 25 the feed for 24 hours. The H₂S addition rate was sufficient to give approximately 250 ppm sulfur on a feed weight basis. Thereafter, the H₂S was removed and the catalysts were only contacted with clean feed and clean 30 hydrogen. The results are shown in Table III.

TABLE III

35	Catalyst	Time	ppm Sulfur (Feed Wt. Basis)	Yield (LV%)	RON (Calc.)
35	0.3% Pt on Mordenite	0-24	0	95.3	79.9
		24-48	234	86.5	80.2
		48-72	0	95.2	79.7
40	0.3% Pt on Beta Zeolite	0-24	0	97.5	79.7
		24-48	234	95.3	79.7
		48-72	0	97.3	79.2

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05 The conventional platinum on mordenite catalyst suffers a relatively larger loss in liquid yield than platinum on Beta zeolite catalyst. Additionally, when sulfur is withdrawn, both catalysts recover from sulfur poisoning, but the liquid yield from the platinum/Beta zeolite is relatively higher.

Example 7

10 This example demonstrates a preferred method of platinum loading which includes a steam calcination step after platinum impregnation.

15 Beta zeolite was prepared in the manner set forth in Example 1. Platinum was impregnated onto the Beta zeolite to achieve 0.3 wt.% and then samples of the catalyst were subjected to different calcination treatments. The different catalysts were then used to isomerize the feed of Table I at 200 psig, 1 LHSV, 6 H₂/HC, for 20 hours. The results are shown in Table IV.

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TABLE IV

Catalyst	Calcination Conditions	Reaction Temperature	Isomerate Properties			
			i/nC ₅	2,2-DMB/nC ₆	Yield (LV%)	RON (Calc.)
25 0.3% Pt/ Beta	500°F in Air	520°F	1.7	0.6	97.4	79.5
0.3% Pt/ Beta	500°F in 50% Steam/ 50% Air	515°F	1.8	1.0	97.8	79.9
30 0.3% Pt/ Beta	600°F in 50% Steam/ 50% Air	520°F	2.1	0.9	97.0	80.1

35 When the catalysts were calcined in a steam/air mixture, the product ratios of iso to normal C₅ hydrocarbons (i/nC₅) and 2,2-dimethylbutane to normal C₆ hydrocarbons (2,2-DMB/nC₆) increased, which in turn increased the product RON. Additionally, this increase was achieved at

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a lower reaction temperature which indicated an increase in catalyst activity.

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The embodiments of this invention which are exemplified above are intended solely as illustrations of the invention. They should not be interpreted as limiting the scope of the invention to just those features which are shown or disclosed. As those familiar with this area of research will appreciate, there are numerous variations of the invention as defined in the following claims which may have not been exemplified but which will achieve equivalent results.

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WHAT IS CLAIMED IS:

- 05 1. A process for isomerizing C₄ to C₇ hydrocarbons, comprising contacting a catalyst, comprising at least one Group VIII metal and Beta zeolite, with a feed having C₄ to C₇ hydrocarbons under isomerization conditions.
- 10 2. A process in accordance with Claim 1 wherein the catalyst has been calcined in a steam/air mixture at an elevated temperature after impregnation of the Group VIII metal.
- 15 3. A process in accordance with Claim 1 wherein the Group VIII metal is platinum.
4. A process in accordance with Claim 3 wherein the catalyst comprises between 0.05% and 5% platinum.
- 20 5. A process in accordance with Claim 3 wherein the catalyst comprises between 0.01% and 2.5% platinum.
- 25 6. A process in accordance with Claim 1 wherein the feed comprises normal and singly branched C₅ and C₆ hydrocarbons.
- 30 7. A process in accordance with Claim 6 wherein the temperature is between 200°F and 1000°F, the pressure is between 50 and 1000 psig, the LHSV is between 1 and 10, and a H₂/HC ratio between 0.5 and 10.
- 35 8. A process for alkane isomerization, comprising: contacting a feed containing normal and singly branched C₅ and C₆ hydrocarbons with a catalyst, which comprises Beta zeolite and between 0.1% and 1.0% platinum, at a temperature between 400°F and 600°F, at a pressure between 100 and 500 psig, a H₂/HC ratio between 1 and 8, and between 1 and 4 LHSV.

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9. A process for alkane isomerization, comprising: contacting a feed containing normal and singly branched C₅ and C₆ hydrocarbons and less than 0.1 ppm sulfur with a catalyst, which comprises Beta zeolite and between 0.1% and 1.0% platinum, at a temperature between 400°F and 600°F, at a pressure between 100 and 500 psig, a H₂/HC ratio between 1 and 8, and between 1 and 4 LHSV.

10. 10. An alkane isomerization catalyst for isomerizing C₄ to C₇ hydrocarbons, comprising a Group VIII metal and Beta zeolite wherein the catalyst has been calcined in a steam/air mixture at an elevated temperature after 15 impregnation with the Group VIII metal.

11. 11. A catalyst in accordance with Claim 10 wherein the Group VIII metal is platinum.

20 12. 12. A catalyst in accordance with Claim 11 wherein the catalyst comprises between 0.05% and 5% platinum.

13. 13. A catalyst in accordance with Claim 12 wherein the catalyst comprises between 0.1% and 2.5% platinum.

25 14. 14. A catalyst in accordance with Claim 10 wherein the Beta zeolite has a silica to alumina ratio between 15 and 100.

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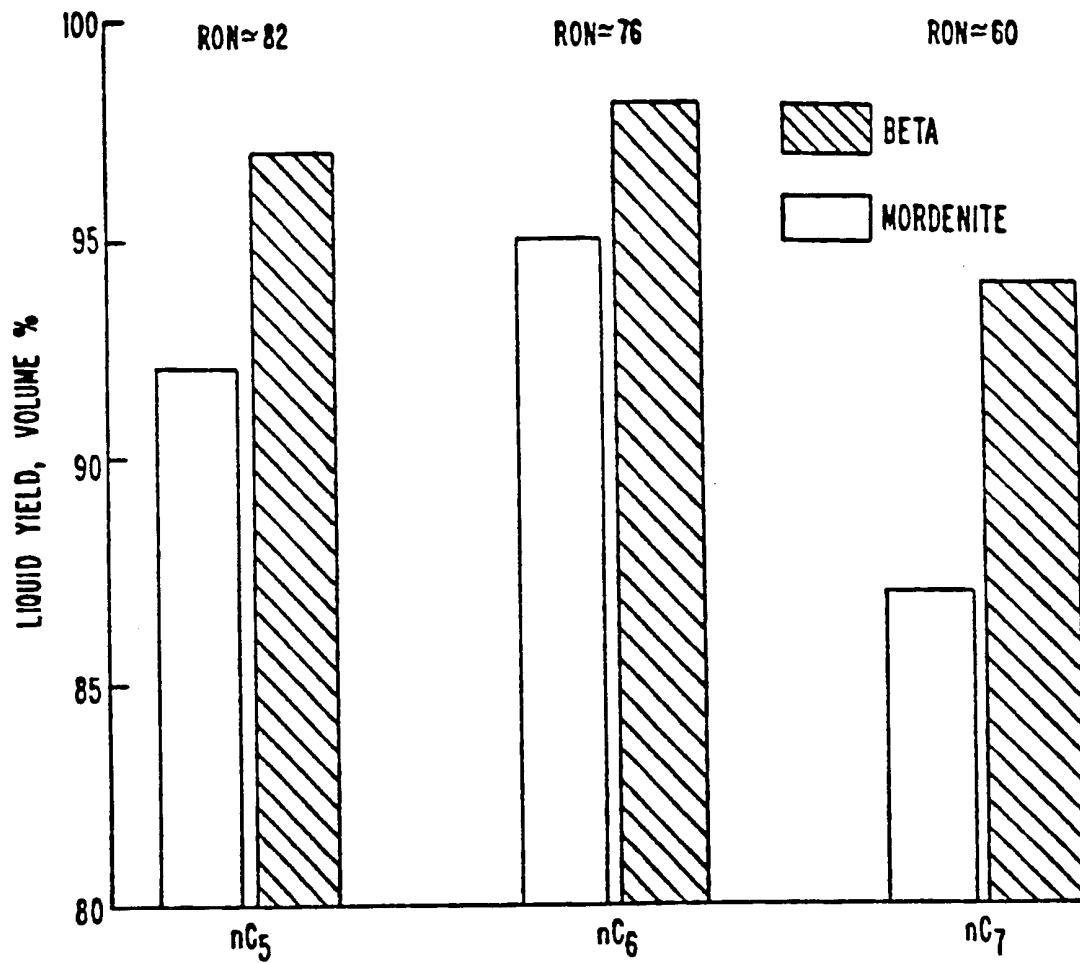


FIG. I.

SUBSTITUTE SHEET

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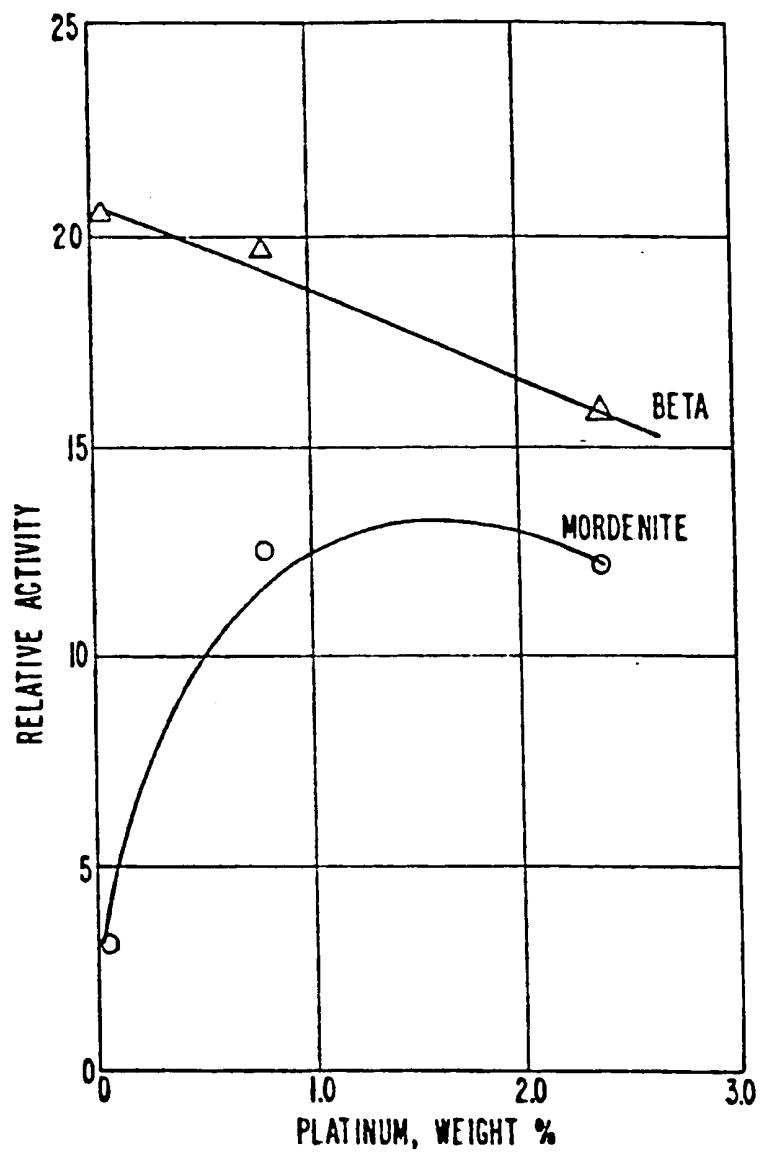


FIG. 2.

INTERNATIONAL SEARCH REPORT

International Application No.

PCT/US90/03655

I. CLASSIFICATION OF SUBJECT MATTER (if several classification symbols apply, indicate all) ¹³

According to International Patent Classification (IPC) or to both National Classification and IPC

IPC (5) : C07C 5/13,5/33; B01J 29/32
U.S. Cl : 585/739,751; 502/66

II. FIELDS SEARCHED

Minimum Documentation Searched ¹⁴

Classification System	Classification Symbols
U.S.	585/739,751; 502/66
Documentation Searched other than Minimum Documentation to the Extent that such Documents are Included in the Fields Searched ¹⁵	

III. DOCUMENTS CONSIDERED TO BE RELEVANT ¹⁴

Category ¹⁶	Citation of Document, ¹⁷ with indication, where appropriate, of the relevant passages ¹⁸	Relevant to Claim No. ¹⁹
Y	US,A 4,642,226 (CALVERT ET AL) 10 February 1987 See column 4, lines 43-68 and columns 5 & 6.	ALL
Y	US,A 4,554,145 (RUBIN) 19 November 1985 See column 6, lines 5-15.	ALL
Y	US,A 4,518,485 (LAPIERRE ET AL) 21 May 1985 See the claims.	ALL
Y	US,A 4,419,220 (LAPIERRE ET AL) 06 December 1983 See column 6, lines 51-68, the paragraph bridging columns 7 & 8, the examples, & the claims.	ALL

(CON'T)

* Special categories of cited documents: ¹⁵

- "A" document defining the general state of the art which is not considered to be of particular relevance
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"X" document of particular relevance; the claimed invention cannot be considered novel or cannot be considered to involve an inventive step

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"&" document member of the same patent family

IV. CERTIFICATION

Date of the Actual Completion of the International Search ²⁰

30 AUGUST 1990

Date of Mailing of this International Search Report ²¹

06 DEC 1990

International Searching Authority ²²

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Signature of Authorized Officer ²³ PCTC-NI

INTERNATIONAL DIVISION
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III. DOCUMENTS CONSIDERED TO BE RELEVANT (CONTINUED FROM THE SECOND SHEET)

Category *	Citation of Document, ¹⁶ with indication, where appropriate, of the relevant passages ¹⁷	Relevant to Claim No ¹⁸
Y	US,A 4,501,926 (LAPIERRE ET AL) 26 February 1985 See the paragraph bridging columns 6 & 7, the paragraph bridging columns 7 & 8, the examples, and the claims.	ALL
Y	USA, 3,308,069 (WADLINGER ET AL) 07 March 1967.	10-14
Y	US,A 4,554,065 (ALBINSON ET AL) 19 November 1985 See the paragraph bridging columns 2 & 3, the paragraph bridging columns 5 & 6 and the examples.	10-14